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Chapter 10 Chemical Quantities. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. claire_kutchka. mrs. wright. Terms in this set (38) standard temperature and pressure. 0 C and 101.3 kPa. the volume occupied by a mole of any gas at STP (22.4 L) molar volume.

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Section 10.1 - The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance. 6.02×10^{23} is called Avogadro's number. 1 mole = 6.02×10^{23} representative particles

Chapter 10 - Chemical Quantities

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Chapter 10 Chemical Quantities91 SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287-289) 1.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A ... After reading Lesson 10.1, answer the following questions. ... The quantities 1 mol and 22.4 L can be used in conversion factors that change moles to volume and volume to moles at STP.

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