

## Chapter 11 The Mole

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### Chapter 11 The Mole

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Chapter 11 The Mole. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. ecsheldon13. Terms in this set (7) mole. the SI base unit used to measure the amount of a substance (310) Avogadro's Number.  $6.02 \times 10^{23}$  (310) molar mass. the mass in grams of one mole of any pure substance (313)

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mole - the SI base unit used to measure the amount of a substance - the number of representative particles in the molar mass of a substance - a mole of anything contains  $6.02 \times 10^{23}$  representative particles

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Chapter 11- The Mole Chapter 12- Stoichiometry Mr. Caldwell's Classroom: The Mole. SI base unit used to measure the amount of a substance  $602,000,000,000,000,000,000,000$  abbreviated to  $6.02 \times 10^{23}$  also known as Avogadro's number. Size of One Mole. Avogadro's number is an enormous number, which is necessary for counting molecules. ...

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312 Chapter 11 The Mole Converting Number of Representative Particles to Moles Zinc is used as a corrosion-resistant coating on iron and steel. It is also an essential trace element in your diet. Calculate the number of moles that contain  $4.50 \times 10^{24}$  atoms of zinc (Zn). 1. You are given the number of atoms of zinc and must find the equivalent number of moles.

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Chapter 11: The Mole 11.1 Measuring Matter counting particles, Avogadro's Number  $6.02 \times 10^{23}$ , mole, mole  $\leftrightarrow$  particle equations # moles  $\times 6.022 \times 10^{23}$  particles/mole = # of particles # particles  $\times 1 \text{ mole} / 6.022 \times 10^{23}$  particles = # of moles. 11.2 Mass and the Mole mass of a mole, molar mass, mass  $\leftrightarrow$  mole equations # of moles  $\times$  # of grams / mole = # of grams

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the formula with the smallest whole number mole ratio of the elements, may or may not be the same as the actual molecular formula. The molecular formula. specifies the actual number of atoms of each element in one molecule in the compound. naming hydrates. ... chemistry chapter 11 the mole. 32 terms. saramukherjee.

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-moles  $\times 6.02 \times 10^{23} / 1 \text{ mole} = c$  -to calc the number of particles of each element of the compound take  $c$  and multiply it by the number of molecules (Ex  $2 \text{ Al}$ ) -Then to find the mass in grams of one unit of the compound take the molar mass and divide it by Avogadro's number.

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In the rest of Chapter 11 we will use the mole to calculate the outcomes of reactions, how much of one substance reacts with another, how many molecules a mass of a substance contains, what are chemical reactions and how do we account for the number of molecules of one kind that react with another and how many different products molecules the reaction produces.

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Chapter 11/4b AP Sample Problem Answers. 15. C. 26. C. 27. C. 36. D. 43. A. 67. B. 71. E. 72. A. 1991 2) a) three points.  $7.2 \text{ g H}_2\text{O} \div 18.0 \text{ g/mol} = 0.40 \text{ mol H}_2\text{O}$ .  $0.40 \text{ mol H}_2\text{O} \times (2 \text{ mol H} / 1 \text{ mol H}_2\text{O}) = 0.80 \text{ mol H}$ .  $7.2 \text{ L CO}_2 \div 22.4 \text{ L/mol} = 0.32 \text{ mol CO}_2$ .  $0.32 \text{ mol CO}_2 \times (1 \text{ mol C} / 1 \text{ mol CO}_2) = 0.32 \text{ mol C}$ . OR

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162 Chemistry: Matter and Change • Chapter 10 Solutions Manual CHAPTER 10 SOLUTIONS MANUAL 10. Explain how a

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mole is similar to a dozen. The mole is a unit for counting  $6.02 \times 10^{23}$  representative particles. The dozen is used to count 12 items. 11. Apply How does a chemist count the number of particles in a given number of moles of a substance?

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Chemistry Matter and Change: Chapter 11 Stoichiometry  
questionStoichiometry answerstudy of quantitative relationships between the amounts of reactants used and the amounts of products formed in a chemical. Samples; ... Mole Ratios: 4 moles Al / 3 moles O<sub>2</sub> or 3 moles O<sub>2</sub> / 4 moles Al 4 moles Al / 2 moles 2Al<sub>2</sub>O<sub>3</sub> or 2 moles Al<sub>2</sub>O<sub>3</sub>/ 4 ...

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