

Chemical Quantities Chapter 10

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Chemical Quantities Chapter 10

Section 10.1 - The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02×10^{23} representative particles of that substance. 6.02×10^{23} is called Avogadro's number. 1 mole = 6.02×10^{23} representative particles

Chapter 10 - Chemical Quantities

CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

CHAPTER 10: Chemical Quantities

Chapter 10 Chemical Quantities. mole. Avogadro's number. representative particle. molar mass. the SI unit for measuring the amount of a substance. number of representative particles in a mole, 6.02×10^{23} . refers to the species present in a substance: usually atoms, m... the mass of one mole of a pure substance.

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Chapter 10 Chemical Quantities. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. claire_kutchka. mrs. wright. Terms in this set (38) standard temperature and pressure. 0 C and 101.3 kPa. the volume occupied by a mole of any gas at STP (22.4 L) molar volume.

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Chapter 10 chemistry chemical quantities. Mole. Avgadro's number. Representative particle. Molar mass. the amount of a substance that contains 6.02×10^{23} representat.... the number of representative particles contained in one mole o.... the smallest unit into which a substance can be broken down wi....

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Chapter 10 "Chemical Quantities" ... % comp from the chemical formula If we know formula, assume you have 1 mole, Subscripts used to calculate mass of each element in 1 mole of cmpd sum of masses is molar mass % Composition Examples % composition as

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Chapter 10 Chemical Quantities. mole. Avogadro's number. representative particle. molar mass. the SI unit for measuring the amount of a substance. number of representative particles in a mole, 6.02×10^{23} . refers to the species present in a substance: usually atoms, m... the mass of one mole of

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a pure substance.

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Chapter 10 Chemical Quantities 91. SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287–289)

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296)

Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A Measurement of Matter Essential Understanding The mole represents a large number of very small particles. Reading Strategy Frayer Model The Frayer Model is a vocabulary development tool. The center of the

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10.3 Percent Composition and Chemical Formulas

Chapter 10 "Chemical Quantities" Vocab Chapter 10 (Chemical Quantities) Test Study Guide The mole is the SI unit used to measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly Chapter 10 Chemical Quantities Test - e13components.com

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